

#### 4. MATERIAL BALANCE

##### Material Balance:

Basis: Per hour of operation

$$\begin{aligned}\text{Amount of cumene to be obtained} &= 1 \text{ M ton of cumene per annum.} \\ &= 10^6/330 \text{ tons per day of cumene.} \\ &= 10^6/(330 \times 24) \text{ tons of cumene per hr.} \\ &= 126.26 \times 10^3 \text{ kg of cumene per hr.} \\ &= (126.26 \times 10^3)/120.19 \text{ kmoles of} \\ &\quad \text{cumene per hr.} \\ &= 1050.50 \text{ Kgmole/hr}\end{aligned}$$

Assuming 97% conversion and 2% loss.

$$\begin{aligned}\therefore \text{Cumene required} &= 1050.50/.98 = 1071.94 \text{ Kgmole/hr} \\ &= 128836.32 \text{ Kg/hr}\end{aligned}$$

Hence 128836.32 kg of cumene is required to be produced per hr.

$$\begin{aligned}\text{Propylene required} &= 1071.94/.97 = 1105.09 \text{ Kgmole} \\ &= 1105.09 \times 42 \text{ Kg/hr of propylene} \\ &= 46413.78 \text{ Kg/hr of propylene}\end{aligned}$$

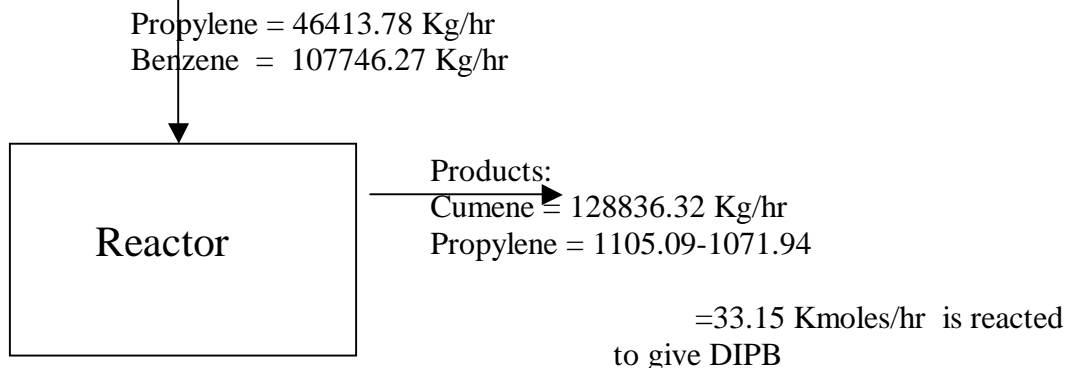
$$\begin{aligned}\text{Assuming benzene required is 25\% extra} \\ &= 1105.09 \times 1.25 \text{ Kmoles of benzene} \\ &= 1381.3625 \text{ Kgmole/hr} \\ &= 107746.27 \text{ Kg/hr}\end{aligned}$$

Propane acts as an inert in the whole process . It is used for quenching purpose in the reactor. It does not take part in the chemical reaction . Also

It is inevitably associated with the propylene as an impurity as their molecular weight is very close. We assume propylene to propane ratio as 3:1.

Being an inert we are neglecting propane balance in the material balance to avoid complexity.

1.) Material balance around reactor :



$$\begin{aligned} \text{Benzene required to give DIPB} \\ &= 33.15/2 \text{ kmoles/hr} \\ &= 16.575 \text{ kmoles/hr} \end{aligned}$$

$$\begin{aligned} \text{DIPB} \\ \text{produced} &= 16.575 \times 162 = 2685.15 \\ &\text{Kg/hr} \end{aligned}$$

$$\begin{aligned} \text{Benzene in product} &= 1381.3625 - 1071.94 - 16.575 \\ &= 292.85 \text{ kmoles/hr} \\ &= 22820.85 \text{ kg/hr} \end{aligned}$$

$$\text{Input} = 46413.78 + 107746.27 = 154160.05 \text{ Kg/hr}$$

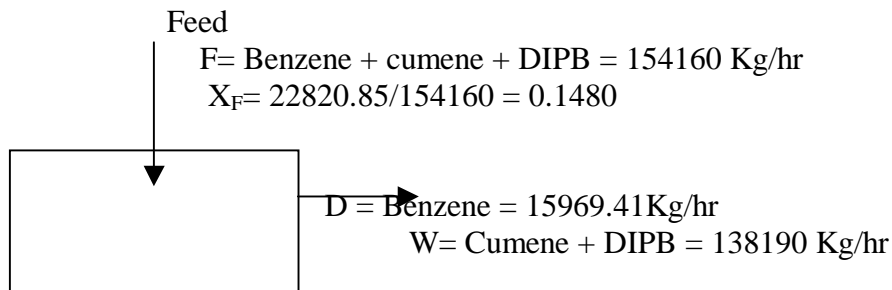
$$\text{Output} = 128836.32 + 2685.15 + 22820.85 = 154342 \text{ Kg/hr}$$

$$\text{Input} = \text{output}$$

## 2) Distillation column 1: ( Depropanasing column )

Assuming almost all the propane is removed in depropanising column and sent to reactor for quenching. Hence material balance for depropanasing column is not considered.

## 3.) Distillation columne 1: (Benzene column)



$$F = D + W$$

$$154160 = D + W$$

$$F X_F = D X_D + W X_w$$

$$\text{Taking } X_F = 0.9999$$

$$X_D = 0.05$$

$$154160 \times 0.1480 = D \times 0.9999 + W \times 0.05$$

$$3023.5 = .9999 D + (20374 - D) \times 0.05$$

$$D = 15969.41 \text{ Kg/hr}$$

$$= \text{Benzene}$$

$$W = 154160 - 15969.41$$

$$= 138190.5 \text{ Kg/hr}$$

$$= \text{cumene} + \text{DIPB}$$

Input = 154160 kg/hr

Output = 15969.41 + 138190.5  
= 154160 kg/hr

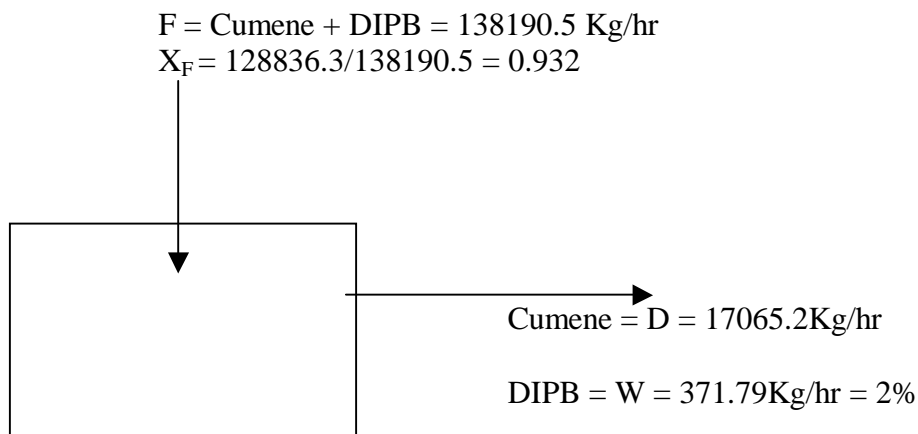
Input = Output

Assuming all the Benzene present in benzene column is recycled to the feed . Hence considering negligible amount of benzene to be part of residue. This will avoid the complexity of multicomponent distillation in Cumene column.

Therefore amount of benzene recycled = 15969.5 Kg/hr.

Therefore feed actually given to the system = 154160 + 15969.5  
= 170129.5 Kg/hr

4.).Distillation column 3: (Cumene column)



$$F = D + W$$

$$138190.6 = D + W$$

$$FX_F = DX_D + WX_W$$

Taking  $X_D = 0.995$

$$X_W = 0.01$$

$$138190.5 \times 0.932 = D \times 0.995 + W \times 0.01$$

$$128793.54 = 0.995D + (138190.5 - D) \times 0.01$$

$$D = 129051 \text{ kg/hr}$$

$$W = 138190.5 - 129051 \\ = 9139.5 \text{ Kg/hr}$$

Input = 138190.5 Kg/hr

$$\begin{aligned}\text{Output} &= 129051 + 9139.5 \\ &= 138190.5 \text{ Kg/hr.}\end{aligned}$$

$$\text{Input} = \text{output}$$